



UTAH GOVERNOR'S OFFICE OF
ENERGY DEVELOPMENT

Helium: An Important Natural Resource

Grade/Subject: Chemistry

Standard CHEM.3.5 Develop solutions related to the management, conservation, and utilization of mineral resources (matter). Define the problem, identify criteria and constraints, develop possible solutions using models, analyze data to make improvements from iteratively testing solutions, and optimize a solution. (PS1.B, ESS3.A, ETS1.A)

Lesson Performance Expectations:

- Students will develop solutions related to well drilling of helium, its conservation and utilization.
- Students will focus on the geochemistry associated with nuclear decay that produces helium in granitic rocks and why drilling is used to recover the gas. A granitic rock is a common, coarse-grained, light-colored, hard igneous rock consisting chiefly of quartz used in monuments and for building and commonly referred to as granite.

Materials: Students need a computer for each pair or individual student.

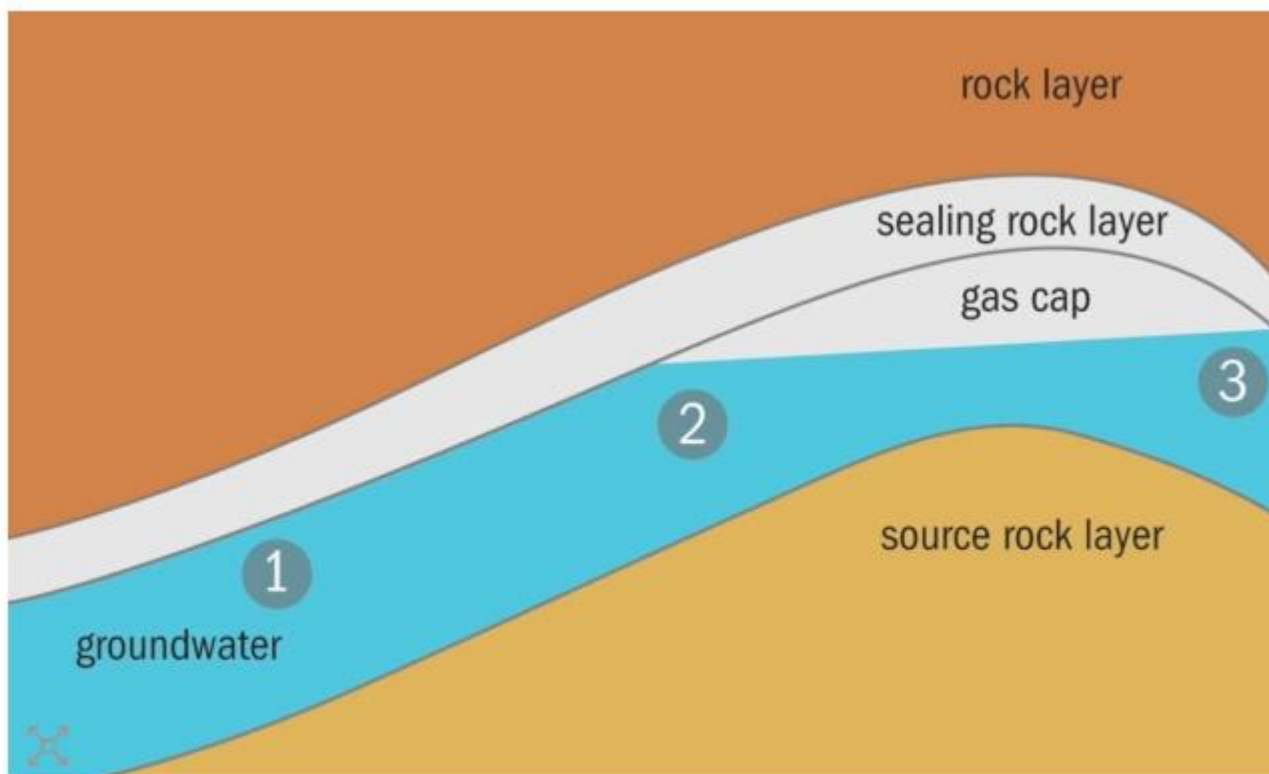
Time: 50 minutes

Teacher Background Information:

- Watch this video about helium <https://www.youtube.com/watch?v=gLCfFugluog&t=9s>
- Basic Geology
 - Helium is an element, which means that it cannot be manufactured under normal conditions.
 - Commercial concentrations of helium are formed under unusual geologic conditions, often in conjunction with natural gas.
 - Helium often forms in Paleozoic-age rocks from radioactive decay of uranium and thorium.
 - Following faults, fractures, and igneous intrusive, helium migrates upward into sedimentary rocks. Igneous intrusions form when lava flows cool and solidify before it reaching the Earth's surface.
 - The atomic radius of helium is so small that it easily moves upward through most of the Earth's layers. Helium can be trapped by non-porous cap-rock such as halite (rock salt) or anhydrite.
 - Igneous rock is one of the three main rock types, the others being sedimentary and metamorphic. Igneous rock is formed through the cooling and solidification of magma or lava.
- Helium Recovery
 - Helium is typically recovered during natural gas extraction and requires large gas fields to produce substantial quantities.
 - Very little helium is present in Earth's atmosphere. It is such a light element that Earth's gravity cannot hold it. When present at the Earth's surface, unconfined helium immediately begins rising until it escapes the planet. That's why party balloons rise!
- Conservation
 - In recent years, shortages of helium have resulted in increased prices. This has led to an interest in drilling helium-only wells.
 - The total terrestrial inventory of helium is estimated to be 17,000 trillion scf (470 trillion scm)
 - Most of this supply is in Earth's atmosphere at a concentration of only 5 ppm.
 - Atmospheric helium is in dynamic equilibrium between the gain of helium diffusing from Earth's crust (as

a product of radioactive decay and losses of helium into space.

- Helium Use
 - Helium is most commonly associated with balloons but industrial uses include military, aerospace and medical applications.
 - Because it is chemically inert and is a liquid at temperatures near absolute zero it is used to cool the magnets in MRI machines and in air-to-air missile guidance systems.
 - It is a carrier gas for gas chromatography and mass spectrometry.
 - Helium is also used for welding, medical imaging, and making semiconductors
- Utah Helium R
 - Utah is home to the first well in the United States in which helium extraction was the primary purpose.
 - The well is located in Grand County on land managed by the Bureau of Land Management (BLM).
 - This area is known as the Harley Dome and contains a high concentration of helium. Designated a federal helium reserve in 1934, it contains a mixture of nitrogen and up to 7% helium. Although this well was plugged in 2019 additional development is occurring within the state.



Helium migration: Once released from the source rock, helium and nitrogen can interact with groundwater (1), which carries the dissolved gases as it ascends. When the groundwater contacts a pre-existing gas cap containing methane or carbon dioxide (2), the nitrogen and helium partition out of the groundwater into the gas cap (3).

[source](#)



IACX started producing on-purpose helium in 2013 at the Harley Dome field in Utah.

- The Harley Dome Field is currently producing helium ([Source](#))

Links:

<https://www.youtube.com/watch?v=gLCfFugluog&t=9s>

<https://physicsworld.com/a/on-the-hunt-for-helium/>

<https://cen.acs.org/articles/95/i30/helium-way.html>

<https://www.npr.org/2019/11/01/775554343/the-world-is-constantly-running-out-of-helium-heres-why-it-matters>

Student Background Knowledge:

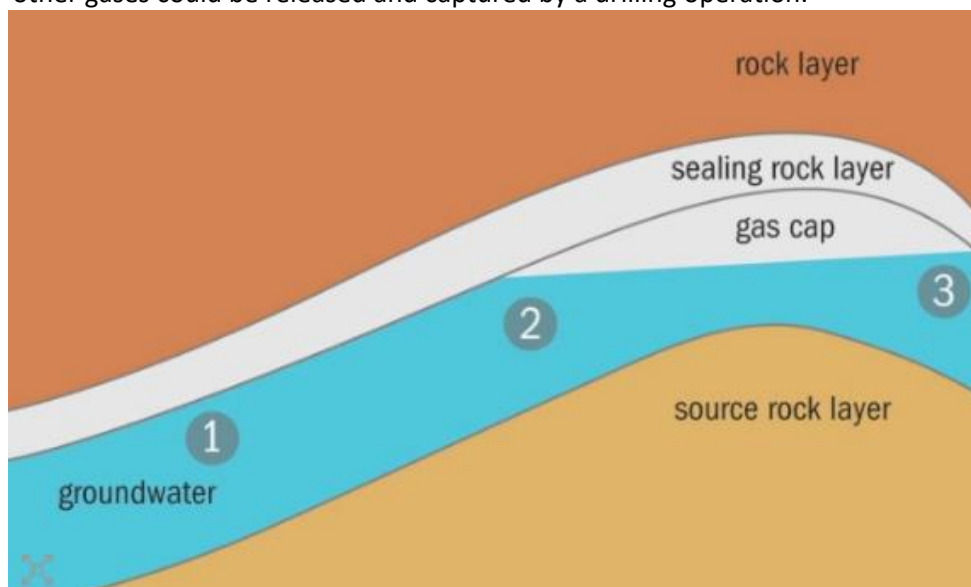
- Students should understand the particulate nature of matter and be know the differences between atoms.
- Students should have a background in radioactive decay.

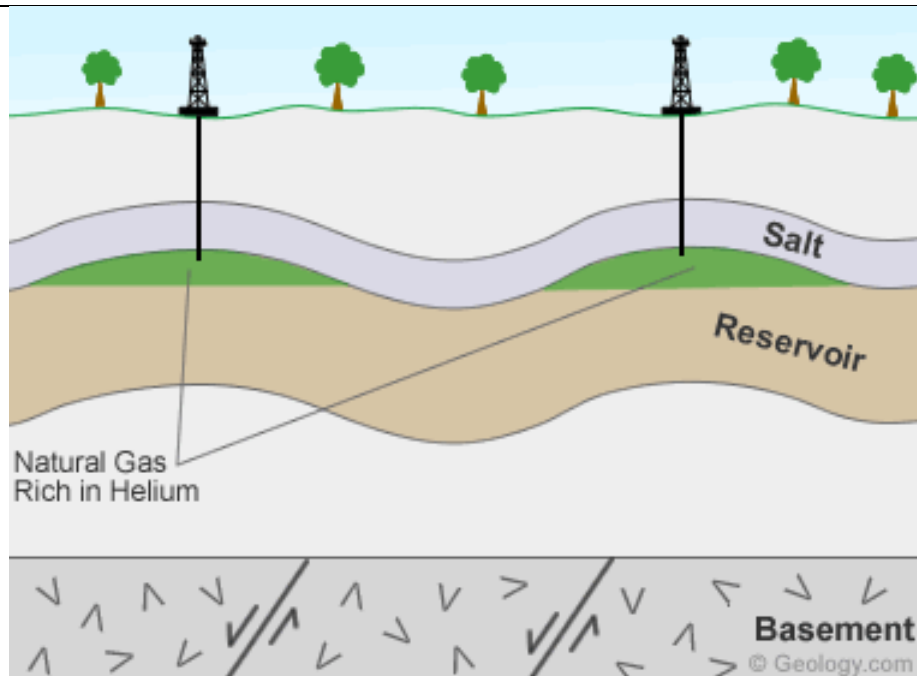
Teacher Step by Step:

1. **Introduce *Phenomenon*:** Bring a helium-filled balloon to class or use a picture of floating balloons.
2. Ask students what they notice about the balloon or balloons pictured on the student sheet.

Assessment of Student Learning.

The diagram below demonstrates a gas cap trapped between source rock and a sealing layer. Helium gas along with other gases could be released and captured by a drilling operation.





1. Most geologists believe that the majority of **helium** in natural gas derives from a specific kind of source rock. What characteristic(s) must this source rock layer have in order to procure helium? Choose all that apply.
 - a. It is formed from large crystals.
 - b. It is porous and allows gases to flow through.*
 - c. It contains radioactive atoms that decay into helium.*
 - d. It contains a solid form of helium that evaporates into gas.

2. What characteristic(s) does the sealing rock layer have? Choose all that apply.
 - a. Helium gas cannot pass through it.*
 - b. It is less dense than the surrounding rocks.
 - c. It was formed after the gas cap was created.
 - d. It is formed from rock containing large crystals.

3. What factors affect the location of the gas cap in the rock formation? Choose all that apply.
 - a. Density*
 - b. Porosity*
 - c. Crystal type
 - d. Temperature
 - e. Regional geology*

4. What characteristics of helium make it useful in many applications? Choose all that apply.
 - a. It is not flammable.*
 - b. It has a low density.*
 - c. It is non-reactive with other elements.*
 - d. It has the ability to “carry” other gases.*
 - e. It turns to liquid at extremely low temperatures.*
 - f. It can be found in many locations and is abundant in nature.

Extension of lesson and Career Connections:

1. Have students design and perform an experiment to measure the density of helium, using balloons as a source of Helium. Check local stores for these balloons.
2. Have students research and report on how helium is used in various applications.